

## *Gpb Chemistry 3 Answer Key 1401*







### Gpb Chemistry 3 Answer Key

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Review Sheet: Unit 3 Name \_\_\_\_\_ CHEMISTRY: A Study of Matter © 2004, GPB 3.23a KEY I. Short Answer and Fill in the Blanks: 1. \_\_\_\_\_ are a familiar idea used to ...

#### I. Short Answer and Fill in the Blanks

The answer you get will be equal to the number of bonding electrons in the ... Chemistry Standard Set 2. Key Vocabulary ... CHEMISTRY: A Study of Matter© 2004, GPB, 5.1.

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Chemistry Episode 901 Answers This first episode is an introduction to the science of chemistry. ourselves and our world better, and in our desire to further Note Taking Guide: Episode 902 Name - Mole Cafe

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3. Choose the coefficients which, in order, would correctly balance the following equation:  $\text{HgO} + \text{O}_2 \rightarrow \text{Hg} + \text{O}_2$  a. 1, 1, 1 b. 2, 1, 1 c. 2, 2, 1 d. 2, 2, 2 4. When balancing a chemical equation, one is allowed to adjust the \_\_\_\_\_. a. formulas of the reactants. b. formulas of the products. c. coefficients d. both A and B

### Chemistry Quiz: Episode 604 Name KEY - PC\|MAC

$3 \times 6 \text{ mol F} + 2 \times 4 \text{ mol AlF}_3 = 30 \text{ mol F}$  2 b. \_\_\_\_\_ moles of  $\text{AlF}_3$  will react with 0.6 moles of  $\text{O}_2$ . ? mol  $\text{AlF}_3 = 0.6 \text{ mol O}_2 \times 4 \text{ mol AlF}_3 / 3 \text{ mol O}_2 = 0.8 \text{ mol AlF}_3$  4. \_\_\_\_  $\text{C}_3\text{H}_8 +$  \_\_\_\_  $\text{O}_2 \rightarrow$  \_\_\_\_  $\text{CO}_2 +$  \_\_\_\_  $\text{H}_2\text{O}$  a. How many moles of oxygen react with 11 moles of  $\text{C}_3\text{H}_8$ ? ? mol  $\text{O}_2 = 11 \text{ mol C}_3\text{H}_8 \times 5 \text{ mol O}_2 / 1 \text{ mol C}_3\text{H}_8 = 55 \text{ mol O}_2$  b. How many moles ...

### CHEMISTRY: A Study of Matter - Weebly

$\text{PCl}_3$  is a compound used to manufacture pesticides. A reaction requires that 96.7 g of  $\text{PCl}_3$  be raised from 31.7 °C to 69.2°C. How much energy will this require given that the specific heat of  $\text{PCl}_3$  is 0.874 J/g °C? 10. A quantity of water is heated from 25.0 °C to 36.4 °C by absorbing 325 J of heat energy. What is the mass of the water?  $325 \text{ J} = (m \dots$

### 13-06a,b,c Heat and Heat Calculations wkst-Key

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### Gpb Chemistry Answer Key 801 - lionandcompass.com

= 3.7!1024 atoms Al 2. Convert 5.3 x 10<sup>25</sup> molecules of CO<sub>2</sub> to moles. ? moles CO<sub>2</sub> = 5.3!10<sup>25</sup> molecules CO<sub>2</sub>! 1 mole CO<sub>2</sub> 6.02!10<sup>23</sup> molecules CO<sub>2</sub> = 88 mol CO<sub>2</sub> 3. How many formula units of sodium acetate are in 0.87 moles of sodium acetate? ? f.u. NaC<sub>2</sub>H<sub>3</sub>O<sub>2</sub> = 0.87 mol NaC<sub>2</sub>H<sub>3</sub>O<sub>2</sub>! 6.02!10<sup>23</sup> f.u. NaC<sub>2</sub>H<sub>3</sub>O<sub>2</sub> 1 mol NaC<sub>2</sub>H<sub>3</sub>O<sub>2</sub> = 5.2!10<sup>3</sup> f ...

### 2 3 2 3 2 1 2H<sub>3</sub>O<sub>2</sub> 3 2 ? f.u. NaCl = 3.55 mol NaCl 6.02 H ...

Given a mass of 24 grams and a volume of 3 milliliters, calculate the density.  $D = \frac{m}{V} = \frac{24 \text{ g}}{3 \text{ mL}} = 8 \text{ g/mL}$  D What is the mass of 32 milliliters of water?  $m = D \cdot V = 1.0 \text{ g/mL} \cdot 32 \text{ mL} = 32 \text{ g}$  Given that the density of iron is 7.9 grams per centimeters cubed, what would be the volume of a 3.5 gram piece of iron?  $V = \frac{m}{D} = \frac{3.5 \text{ g}}{7.9 \text{ g/cm}^3} = 0.44 \text{ cm}^3$

### CHEMISTRY: A Study of Matter - Thomas County Schools

at a pressure of 2 atm, and 1 L of hydrogen gas at a pressure of 3 atm. All 3 samples are at room temperature. If you transfer the oxygen and nitrogen to the container occupied by the hydrogen, the pressure exerted by the oxygen in the final mixture will be \_\_\_\_\_. The pressure exerted by the mixture will be \_\_\_\_\_.

### FILL IN THE BLANKS and SHORT ANSWER

CHEMISTRY: A study of Matter 0 2004, GPB 13.2 When heat (Q) is absorbed by a system, part of it (C) goes into storage as energy and part of it is used to make the molecules move around, raising the (At). Why does sand get hotter in the day and colder at night than the water? CHEMISTRY: A study of Matter e 2004, GPB 13.1

### www.cbsd.org

Show all work and express your answers in scientific notation. 1. 1000 mHow many milligrams are in ten grams? ? mg = 10g ! 1000 mg 1g = 1!10<sup>4</sup> mg 2. A metal bar has a length of 3.0 cm, a width of 3.0 cm, and a height of 2.0 cm. What volume of water, in liters, would the bar displace? 4.0 cm ! 3.0 cm ! 2.0 cm = 24 cm<sup>3</sup>? L = 24 cm<sup>3</sup>! 1 mL 1 cm<sup>3</sup>! 1 L ...

### Show all work and express your answers in scientific notation.

Chemistry & Physics consists of two series teaching high school college preparatory chemistry and physics. Chemistry: A Study of Matter and Physics Fundamentals provides instructional content delivered through thirty-minute episodes and integrated classroom materials.

### Chemistry & Physics | GPB

4Al + 3O<sub>2</sub> 2Al<sub>2</sub>O<sub>3</sub> 2. C<sub>3</sub>H<sub>8</sub> + 5O<sub>2</sub> 3CO<sub>2</sub> + 4H<sub>2</sub>O 3. Al(NO<sub>3</sub>)<sub>3</sub> + 3NaOH Al(OH)<sub>3</sub> + 3NaNO<sub>3</sub> chemical equation reactants products coefficients balance Conservation equal created destroyed coefficients formulas / subscripts

### Worksheet: Balancing Equations Name KEY I. Fill in the ...

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3. What is the molarity of a solution of HNO<sub>3</sub> that contains 12.6 grams HNO<sub>3</sub> in 1.0 L of solution? ? mol HNO<sub>3</sub> = 12.6 g HNO<sub>3</sub> x 1 mol HNO<sub>3</sub> 63.0 g HNO<sub>3</sub> = 0.200 mol HNO<sub>3</sub> M = 0.200 mol HNO<sub>3</sub> 1.0 L = 0.200 M 4. How many grams of potassium nitrate are required to prepare 0.250 L of a 0.700 M solution? 0.700 M = moles of solute 0.250 L moles of ...

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